

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $10^{-14}$  (B)  $10^{-7}$  (C)  $10^{-1}$  (D)  $10^{-1}$

(A)  $10^{-14}$  (B)  $10^{-7}$  (C)  $10^{-1}$  (D)  $10^{-1}$   
 (E)  $10^{-14}$  (F)  $10^{-7}$  (G)  $10^{-1}$  (H)  $10^{-1}$  (I)  $10^{-14}$  (J)  $10^{-7}$  (K)  $10^{-1}$  (L)  $10^{-1}$

2. In a 0.1 M acetic acid solution, the concentration of acetate ions is 0.01 M. The pH of the solution is (correct)
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0 (F) 6.0 (G) 7.0 (H) 8.0 (I) 9.0 (J) 10.0

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

3. The pH of a 0.1 M acetic acid solution is 3.0. The concentration of acetate ions is (correct)
- (A) 0.01 M (B) 0.02 M (C) 0.03 M (D) 0.04 M (E) 0.05 M (F) 0.06 M (G) 0.07 M (H) 0.08 M (I) 0.09 M (J) 0.10 M

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

4. Which of the following is a weak acid? (correct)
- (A) HCl (B) HNO<sub>3</sub> (C) H<sub>2</sub>SO<sub>4</sub> (D) H<sub>3</sub>PO<sub>4</sub> (E) H<sub>2</sub>CO<sub>3</sub> (F) H<sub>2</sub>SO<sub>3</sub> (G) H<sub>2</sub>SO<sub>4</sub> (H) H<sub>2</sub>SO<sub>3</sub> (I) H<sub>2</sub>SO<sub>4</sub> (J) H<sub>2</sub>SO<sub>3</sub>

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

5. A 0.1 M acetic acid solution is mixed with 0.1 M sodium acetate. The pH of the solution is (correct)
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0 (F) 6.0 (G) 7.0 (H) 8.0 (I) 9.0 (J) 10.0

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

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6. The pH of a 0.1 M acetic acid solution is 3.0. The concentration of acetate ions is (correct)
- (A) 0.01 M (B) 0.02 M (C) 0.03 M (D) 0.04 M (E) 0.05 M (F) 0.06 M (G) 0.07 M (H) 0.08 M (I) 0.09 M (J) 0.10 M

(A)  $10^{-1}$  (B)  $10^{-2}$  (C)  $10^{-3}$  (D)  $10^{-4}$  (E)  $10^{-5}$  (F)  $10^{-6}$  (G)  $10^{-7}$  (H)  $10^{-8}$  (I)  $10^{-9}$  (J)  $10^{-10}$

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**Roger Slade, Maureen Slade**



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