

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are the characteristic properties for a weak acid? _____
- (A) pKa and molar concentration (B) pKa and pH
 (C) molar concentration (D) molar concentration and pH

2. Which of the following is CORRECT? In a weak acid solution, the concentration of the undissociated acid is _____
- (A) 10^{-4} mol/L (B) 10^{-5} mol/L
 (C) 10^{-6} mol/L (D) 10^{-7} mol/L

3. In a 0.10 mol/L weak acid solution, the concentration of the undissociated acid is 0.01 mol/L. The pH of the solution is _____
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0

4. In a 0.10 mol/L weak acid solution, the concentration of the undissociated acid is 0.01 mol/L. The pH of the solution is _____
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0

5. Which of the following is CORRECT? In a weak acid solution, the concentration of the undissociated acid is _____
- (A) 10^{-4} mol/L (B) 10^{-5} mol/L (C) 10^{-6} mol/L (D) 10^{-7} mol/L

6. Which of the following is CORRECT? In a weak acid solution, the concentration of the undissociated acid is _____
- (A) 10^{-4} mol/L (B) 10^{-5} mol/L (C) 10^{-6} mol/L (D) 10^{-7} mol/L

7. In a 0.10 mol/L weak acid solution, the concentration of the undissociated acid is 0.01 mol/L. The pH of the solution is _____
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0

8. In a 0.10 mol/L weak acid solution, the concentration of the undissociated acid is 0.01 mol/L. The pH of the solution is _____
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0

9. In a 0.10 mol/L weak acid solution, the concentration of the undissociated acid is 0.01 mol/L. The pH of the solution is _____
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0

Multiple Choice Stoichiometry Test

Tobias Bleicker



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