

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations, pH, and pOH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|---------------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair | 3. pH |
| 4. hydronium ion | 5. acid ionization constant (K_a) | 6. pOH |
| 7. hydroxide ion | 8. weak acid | |

Key Equations

- | | |
|--|--|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pOH} = -\log[\text{OH}^-]$ | 4. $\text{pH} + \text{pOH} = 14$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the _____ of a weak hydroxide ion and a hydronium ion? _____
2. The conjugate base of a weak acid is _____.
3. _____ is the conjugate acid of a weak base.
4. _____ is the conjugate base of a weak acid.
5. _____ is the conjugate acid of a weak base.
6. _____ is the conjugate base of a weak acid.
7. _____ is the conjugate acid of a weak base.
8. _____ is the conjugate base of a weak acid.
9. _____ is the conjugate acid of a weak base.
10. _____ is the conjugate base of a weak acid.
11. _____ is the conjugate acid of a weak base.
12. _____ is the conjugate base of a weak acid.
13. _____ is the conjugate acid of a weak base.
14. _____ is the conjugate base of a weak acid.
15. _____ is the conjugate acid of a weak base.
16. _____ is the conjugate base of a weak acid.
17. _____ is the conjugate acid of a weak base.
18. _____ is the conjugate base of a weak acid.
19. _____ is the conjugate acid of a weak base.
20. _____ is the conjugate base of a weak acid.

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Raffaella Di Napoli



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