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**CODES: A STEP-BY-STEP GUIDE**

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 fe<sup>2+</sup> co co<sup>2+</sup> fe oxidation co co<sup>2+</sup> 2e reduction fe<sup>2+</sup> 2e fe 24 3 ag ni ni<sup>3+</sup> 3 ag oxidation reduction 25 cu<sup>2+</sup> pb pb<sup>2+</sup> cu oxidation reduction 26 o 2

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 $2\text{O}_5$   
 $5\text{CO}_2$   
 $2\text{O}_5$   
 $5\text{CO}_2$

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 $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   
 $\text{Sr}^0$  to  $\text{Sr}^2$  oxidized reducing agent  $\text{O}_2$  to  $\text{O}^2$  reduced  
 $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$   
 $\text{Li}^0$  to  $\text{Li}^1$  oxidized red ag  
 $\text{S}^0$  to  $\text{S}^2$  reduced ox ag  $3\text{C} + 2\text{S} \rightarrow \text{CS}_2$

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