

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (B) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (D) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$ (E) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (F) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. An acid-base reaction is represented by the equation: $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$. The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} .

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3. The reaction between sodium hydroxide and hydrochloric acid is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} .

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4. Which of the following is a correct statement? (correct)

(A) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (B) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (D) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$ (E) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (F) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

5. A solution of sodium hydroxide and hydrochloric acid is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} .

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6. The reaction between sodium hydroxide and hydrochloric acid is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} . The reaction is exothermic. The enthalpy change for the reaction is -57 kJ mol^{-1} .

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