

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)**1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- Mass of an object? **triple beam balance**
 - Volume of an object? **ruler or water displacement method**
 - Length of an object? **Ruler**
 - Density of an object? **graduated cylinder**
- Basic SI Unit? **kg**
 Basic SI Unit? **cm³ or ml**
 Basic SI Unit? **meter**
 Basic SI Unit? **ml**
- b. What is the formula for calculating density? **mass/volume**
- What is the density of a cube that is 35 g and has a volume of 3 cm³? **35/3cm³**
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron.	Plum Pudding Model
Niels Bohr	1913	Electrons move fixed orbits around the atom's nucleus in specific layers, or electron shells . Atoms absorb or give off energy when the electrons move from one shell (energy level) to another.	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus.	Planetary Model

4. Periodic Table

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- An element's **properties** can be predicted from its location in the periodic table.
 - The atomic number represents the number of **protons**.
 - The atomic mass represents the number of **protons and neutrons**.

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IM Harris



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Table of Contents 2nd Semester Final Exam Study Guide Answers

1. Understanding the eBook 2nd Semester Final Exam Study Guide Answers
 - The Rise of Digital Reading 2nd Semester Final Exam Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Final Exam Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Final Exam Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Final Exam Study Guide Answers
 - Personalized Recommendations
 - 2nd Semester Final Exam Study Guide Answers User Reviews and Ratings
 - 2nd Semester Final Exam Study Guide Answers and Bestseller Lists
5. Accessing 2nd Semester Final Exam Study Guide Answers Free and Paid eBooks
 - 2nd Semester Final Exam Study Guide Answers Public Domain eBooks
 - 2nd Semester Final Exam Study Guide Answers eBook Subscription Services
 - 2nd Semester Final Exam Study Guide Answers Budget-Friendly Options
6. Navigating 2nd Semester Final Exam Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 2nd Semester Final Exam Study Guide Answers Compatibility with Devices
 - 2nd Semester Final Exam Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2nd Semester Final Exam Study Guide Answers
 - Highlighting and Note-Taking 2nd Semester Final Exam Study Guide Answers
 - Interactive Elements 2nd Semester Final Exam Study Guide Answers

8. Staying Engaged with 2nd Semester Final Exam Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Final Exam Study Guide Answers
9. Balancing eBooks and Physical Books 2nd Semester Final Exam Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Final Exam Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Final Exam Study Guide Answers
 - Setting Reading Goals 2nd Semester Final Exam Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2nd Semester Final Exam Study Guide Answers
 - Fact-Checking eBook Content of 2nd Semester Final Exam Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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