

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|---|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion (H ₃ O ⁺) | 6. neutral solution |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|--------------------------------------|--------------------------------------|
| 1. $[H^+] = 10^{-pH}$ (see page 5) | 2. $pH = -\log [H^+]$ (see page 5) |
| 3. $[OH^-] = 10^{-pOH}$ (see page 5) | 4. $pOH = -\log [OH^-]$ (see page 5) |

Part A: Completion

Use the appropriate number to indicate the relationship of the conjugate acid-base pair indicated in the second column with the number in the first column.

1. What product is the _____ of a weak hydroxide ion and a hydroxide ion? _____ of the conjugate base of that weak hydroxide ion? _____ of _____.
2. Hydroxide ion and hydronium ions _____ in water.
3. _____ is the conjugate base of the hydroxide ion. _____ is the conjugate _____ of _____.
4. _____ is the conjugate base of a weak acid and _____ is the conjugate acid of a weak base.
5. _____ is the conjugate acid of a weak base and _____ is the conjugate base of a weak acid.
6. _____ is the conjugate acid of a weak base and _____ is the conjugate base of a weak acid.

7. _____
8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

19 2 Hydrogen Ions And Acidity Answer Key

Zhe Xu



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