

# Section 3.1

## Measurements and Their Uncertainty

- OBJECTIVES:

- Determine the number of significant figures in a measurement and in a calculated answer.

# Section 3 1 Measurements And Their Uncertainty

## Practice Problems

**Clemens Wendtner**



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web jun 2 2020 the reactant that is consumed first and limits the amount of product s that can be obtained is the limiting reactant to identify the limiting reactant calculate the number of moles of each reactant present and compare this ratio to the mole ratio of the reactants in the balanced chemical equation

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web determining the limiting reactant in the real world amounts of reactants and products are typically measured by mass or by volume it is first necessary to convert the given quantities of each reactant to moles in order to identify the limiting reactant

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web limiting reactant the substance with the lowest concentration in a chemical reaction its amount determines the amount

of product made from that reaction see also reactant

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web jun 30 2023 the limiting reagent is the one that is totally consumed it limits the reaction from continuing because there is none left to react with the in excess reactant there are two ways to determine the limiting reagent

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web a which chemical is the limiting reactant b how many grams of zns will be formed c how many grams of the excess reactant will remain after the reaction is over 3 which element is in excess when 3 00 grams of mg is ignited in 2 20 grams of pure oxygen what mass is in excess what mass of mgo is formed 4

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web reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated molecular equation chemical equation in which all reactants and products are represented as neutral substances

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web the key to recognizing which reactant is the limiting reagent is based on a mole mass or mass mass calculation whichever reactant gives the lesser amount of product is the limiting reagent what we need to do is determine an amount of one product either moles or mass assuming all of each reactant reacts

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web a crucial skill in evaluating the conditions of a chemical process is to determine which reactant is the limiting reagent

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and which is in excess the key to recognizing which reactant is the limiting reagent is based on a mole mass or mass mass calculation whichever reactant gives the lesser amount of product is the limiting reagent what we

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web the reactant which is in a lesser amount than is required by stoichiometry is the limiting reactant in an alternate method of finding the limiting reagent the amount of product formed by each reactant is calculated the limiting reactant is the reactant from which the minimum amount of product is formed also if we calculate the amount of

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web the limiting reactant or limiting reagent is the reactant that gets consumed first in a chemical reaction and therefore limits how much product can be formed as we saw in example 1 there are many different ways to determine the limiting reactant but they all involve using mole ratios from the balanced chemical equation

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web may 1 2013 the limiting reactant or limiting reagent is the reactant that determines the amount of product that can be formed in a chemical reaction the reaction proceeds until the limiting reactant is completely used up in our example above

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