

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|------------------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 14.1) | 2. $pH = -\log[H_3O^+]$ (Eq. 14.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 14.3) | 4. $pH = -\log[H_3O^+]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to indicate the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction with the conjugate species.

1. When protons are _____, a conjugate hydroxide ion and a hydronium ion are _____ the conjugates of these species produced in the autoionization of water.
2. Hydrochloric acid ionizes from _____ to _____.
3. _____ ionizes to produce _____ and _____.
4. _____ ionizes to form _____ and _____.
5. _____ ionizes to form _____ and _____.
6. _____ ionizes to form _____ and _____.
7. _____ ionizes to form _____ and _____.
8. _____ ionizes to form _____ and _____.

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