

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities (correct)?
 (A) pressure, concentration (B) pressure, pH
 (C) molar fraction (D) dimensional analysis
2. Which of the following is CORRECT? (A) concentration of ions increases as pH increases (B) concentration of ions increases as pH decreases



4. In a 0.1 M solution of acetic acid, CH_3COOH , $[\text{CH}_3\text{COO}^-] = 0.001 \text{ M}$, $[\text{CH}_3\text{COOH}] = 0.099 \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the pH of this solution?



5. In a 0.1 M solution of CO_2 , $[\text{CO}_3^{2-}] = 10^{-6} \text{ M}$, $[\text{HCO}_3^-] = 10^{-3} \text{ M}$, $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the pH of this solution?



6. Which of the following is CORRECT? (A) $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$ (B) $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$ (C) $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$ (D) $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$



7. A solution is made by adding 10 mL of 0.1 M HCl to 10 mL of 0.1 M NaOH . What is the pH of this solution?



8. In a 0.1 M solution of H_2CO_3 , $[\text{H}^+] = 10^{-4} \text{ M}$, $[\text{HCO}_3^-] = 10^{-3} \text{ M}$, $[\text{CO}_3^{2-}] = 10^{-6} \text{ M}$, $[\text{OH}^-] = 10^{-10} \text{ M}$. What is the pH of this solution?



Multiple Choice Stoichiometry Test

Camilla Rothe

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