

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $n = 1000$, $m = 1000$

(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$
 (C) $1 \text{ mol H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

2. An acid-base indicator has a pK_a of 4.5. In a solution of pH 5.5, the indicator is present in the form of the acid. What is the ratio of the concentration of the acid to the concentration of the base? (correct)

(A) 10^{10} (correct) (B) 10^5
 (C) 10^4 (D) 10^3

(E) 10^2 (F) 10^1 (G) 10^0 (H) 10^{-1} (I) 10^{-2}

3. An acid-base indicator has a pK_a of 4.5. In a solution of pH 5.5, the indicator is present in the form of the acid. What is the ratio of the concentration of the acid to the concentration of the base? (correct)

(A) 10^5 (B) 10^4 (C) 10^3 (D) 10^2

4. Which of the following is a dimensional quantity? (correct)

(A) $n = 1000$ (B) $m = 1000$

(C) $10 \text{ g H}_2\text{O}$ (D) $1 \text{ mol H}_2\text{O}$ (E) $1 \text{ g H}_2\text{O}$ (F) $1 \text{ mol H}_2\text{O}$

5. A solution of an acid has a pH of 3.5. The concentration of the acid is 0.1 mol/L. What is the concentration of the base? (correct)

(A) 10^{-3} (B) 10^{-4} (C) 10^{-5} (D) 10^{-6} (E) 10^{-7}

(F) 10^{-8} (G) 10^{-9} (H) 10^{-10} (I) 10^{-11} (J) 10^{-12}

6. An acid-base indicator has a pK_a of 4.5. In a solution of pH 5.5, the indicator is present in the form of the acid. What is the ratio of the concentration of the acid to the concentration of the base? (correct)

(A) 10^{10} (B) 10^5 (C) 10^4 (D) 10^3 (E) 10^2

Multiple Choice Stoichiometry Test

M Walker



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