



is 54.8. If 0.5 mol L^{-1} of $\text{HI}_{(g)}$ is present at equilibrium at 700 K, what are the concentration of $\text{H}_{2(g)}$ and $\text{I}_{2(g)}$ assuming that we initially started with $\text{HI}_{(g)}$ and allowed it to reach equilibrium at 700 K?

Answer:

It is given that equilibrium constant K_c for the reaction



Therefore, at equilibrium, the equilibrium constant K'_c for the reaction



$$[\text{HI}] = 0.5 \text{ mol L}^{-1}$$

Let the concentrations of hydrogen and iodine at equilibrium be $x \text{ mol L}^{-1}$

$$[\text{H}_2] = [\text{I}_2] = x \text{ mol L}^{-1}$$

$$\text{Therefore, } \frac{[\text{H}_2][\text{I}_2]}{[\text{HI}]^2} = K'_c$$

$$\Rightarrow \frac{x \times x}{(0.5)^2} = \frac{1}{54.8}$$

$$\Rightarrow x^2 = \frac{0.25}{54.8}$$

$$\Rightarrow x = 0.06754$$

$$x = 0.068 \text{ mol L}^{-1} \text{ (approximately)}$$

Hence, at equilibrium, $[\text{H}_2] = [\text{I}_2] = 0.068 \text{ mol L}^{-1}$.

Question 7.16:

What is the equilibrium concentration of each of the substances in the equilibrium when the initial concentration of ICl was 0.78 M ?



Answer:

The given reaction is:

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