

## Answer Sheet for Sample Problems for Chemistry Level 1 Final Exam 2013-2014 Study Guide

### Electrons in Atoms – Chapter 5

1. What is the frequency of green light, which has a wavelength of  $4.90 \times 10^{-7} \text{ m}$ ?

$$c = \lambda \nu \quad \lambda = 4.90 \times 10^{-7} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}}$$

$$\underline{6.12 \times 10^{14} \text{ Hz}}$$

2. An x-ray has a wavelength of  $1.15 \times 10^{-10} \text{ m}$ . What is its frequency?

$$c = \lambda \nu \quad \lambda = 1.15 \times 10^{-10} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}}$$

$$\underline{2.61 \times 10^{18} \text{ Hz}}$$

3. Which type of electromagnetic radiation has the most energy: Microwaves, X-Rays, or Visible light? Why?

**Ans: X-Rays, b/c they have the shortest wavelength and highest frequency of the types listed and frequency and energy are directly related.**

4. What is the energy of a  $9.50 \times 10^{13} \text{ Hz}$  radiation?

$$E = h\nu \quad \nu = 9.50 \times 10^{13} \text{ Hz}, h = 6.626 \times 10^{-34} \text{ J s} \quad E = ? \text{ J}$$

$$E = h\nu = (6.626 \times 10^{-34} \text{ J s}) (9.50 \times 10^{13} \text{ s}^{-1}) = \underline{6.29 \times 10^{-19} \text{ J}}$$

5. The element with the electron configuration of  $[\text{Kr}]5s^2 4d^{10} 5p^1$  is indium (In)

6. Barium (Ba)

7. Write out the electron configurations for the following elements: (For the exam make sure you know how to do both the complete and the noble gas short cut for electron configurations)

a. Tc                      **Ans:  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^5 4p^6 5s^2 4d^5$  or  $[\text{Kr}] 5s^2 4d^5$**

b. Aluminum            **Ans:  $1s^2 2s^2 2p^6 3s^2 3p^1$  or  $[\text{Ne}] 3s^2 3p^1$**

8. Why are Cr and Cu exceptions to writing the electron configurations of elements?

**They are exceptions because instead of following the pattern of filling up the d orbitals Cr actually has a half filled s and d orbitals and Cu has a half filled s and completely filled d orbitals. The reason why this occurs is because the electrons in the d-sublevel are more stable with this electron configuration than the expected configuration.**

9. a. Ca

b. As

c. Sn

a. 2    Ca.

b. 5    .As.

c. 4    . Sn .

### Periodic Table and Periodic Trends – Chapter 6

1. Compare & contrast how Mendeleev's periodic table was organized to how Moseley organized it.

# 2nd Semester Exam Review Sheet Answer

**Daniela Niemeyer**



**2nd Semester Exam Review Sheet Answer:**

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**2nd Semester Exam Review Sheet Answer**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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### **Table of Contents 2nd Semester Exam Review Sheet Answer**

1. Understanding the eBook 2nd Semester Exam Review Sheet Answer
  - The Rise of Digital Reading 2nd Semester Exam Review Sheet Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Exam Review Sheet Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2nd Semester Exam Review Sheet Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Exam Review Sheet Answer
  - Personalized Recommendations
  - 2nd Semester Exam Review Sheet Answer User Reviews and Ratings
  - 2nd Semester Exam Review Sheet Answer and Bestseller Lists

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  - 2nd Semester Exam Review Sheet Answer Public Domain eBooks
  - 2nd Semester Exam Review Sheet Answer eBook Subscription Services
  - 2nd Semester Exam Review Sheet Answer Budget-Friendly Options
6. Navigating 2nd Semester Exam Review Sheet Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - 2nd Semester Exam Review Sheet Answer Compatibility with Devices
  - 2nd Semester Exam Review Sheet Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2nd Semester Exam Review Sheet Answer
  - Highlighting and Note-Taking 2nd Semester Exam Review Sheet Answer
  - Interactive Elements 2nd Semester Exam Review Sheet Answer
8. Staying Engaged with 2nd Semester Exam Review Sheet Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2nd Semester Exam Review Sheet Answer
9. Balancing eBooks and Physical Books 2nd Semester Exam Review Sheet Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2nd Semester Exam Review Sheet Answer
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Exam Review Sheet Answer
  - Setting Reading Goals 2nd Semester Exam Review Sheet Answer
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2nd Semester Exam Review Sheet Answer
  - Fact-Checking eBook Content of 2nd Semester Exam Review Sheet Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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