

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from  $K_w$  and  $K_a$  or  $K_b$  and the concentration of a weak acid or base.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |              |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium      | 5. $K_a$                    | 6. $K_b$     |
|                   | 7. weak base                |              |

## Key Equations

- |                           |                                       |
|---------------------------|---------------------------------------|
| 1. $K_w = [H_3O^+][OH^-]$ | 2. $K_a = \frac{[H_3O^+][A^-]}{[HA]}$ |
| 3. $pH = -\log[H_3O^+]$   | 4. $pOH = -\log[OH^-]$                |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Write the number in the space provided.

1.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (acidic hydrogen transfer)
2.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)
3.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)
4.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)
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7.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)
8.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)
9.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)
10.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (the hydroxide ion of the acidic hydrogen transfer)

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# 19 2 Hydrogen Ions And Acidity Answer Key

**CH Cherryholmes**



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