

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H ₃ O ⁺) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|---------------------------------------|---------------------------------------|
| 1. $[H^+] = 10^{-pH}$ (see page 17.1) | 2. $pH = -\log [H^+]$ (see page 17.1) |
| 3. $pH = -\log [H^+]$ (see page 17.1) | 4. $[H^+] = 10^{-pH}$ (see page 17.1) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
2. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
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8. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
9. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
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19 2 Hydrogen Ions And Acidity Work Sheet Answers

Joacim Rocklöv



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