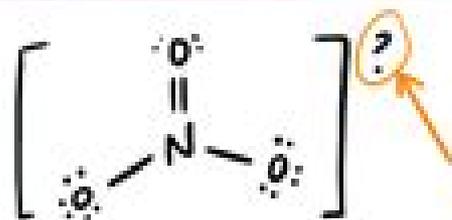


# [ FORMAL CHARGE ]<sup>+/-</sup>

$$F.C. = (\text{valence } e^-) - (\text{nonbonding } e^-) - \frac{1}{2} (\text{bonding } e^-)$$

- \* The amount of electrons in a specific atom when the molecule shares electrons equally.
- \* Only applies to covalent molecules.

## Examples



Overall charge comes from addition of all atomic formal charges!

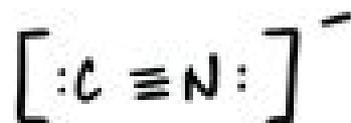
Q: What is the formal charge of N?

A:

$$F.C. = 5 - 0 - \left(\frac{1}{2} \times 8\right)$$

$$F.C. = 5 - 4$$

$$F.C. = +1$$



Q: What is the formal charge of carbon?

$$A: F.C. = 4 - 2 - \left(\frac{1}{2} \times 6\right)$$

$$F.C. = 4 - 2 - 3$$

$$F.C. = -1$$

# Formal Charges

**K Payea**



**Formal Charges:**

## Reviewing **Formal Charges**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Formal Charges**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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