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Facing The Future

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determine the concentrations of solutions of acids and bases through the titration process we are able to identify physical changes to the mixture such as the colour change to indicate the end point of the experiment

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \text{aq} + 2\text{NaOH} \text{aq} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction $\text{NaOH} \text{aq} + \text{C}_6\text{H}_4\text{COOH} \text{aq} \rightarrow \text{C}_6\text{H}_4\text{COONa} \text{aq} + \text{H}_2\text{O} \text{l}$ the point of equivalence is the point of titration when the number of moles of naoh added is equal to the number of moles of khp in the

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