

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. K_a | 6. K_b |
| | 7. weak base | |

Key Equations

- | | |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pOH} = -\log[\text{OH}^-]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the K_a and K_b values listed in Appendix C. Assume that all solutions are aqueous.

1. What products are **1** **2** formed by the autoionization of water? Write the chemical formula of the products of these acids produced in **3** **4** water.
2. Hydrochloric acid has a K_a of **1** **2** . Is it a strong acid? **3** **4** . Write the chemical formula of the conjugate base of **1** **2** . Write the chemical formula of **3** **4** .
3. **1** **2** is the K_a for a weak acid. Write the chemical formula of **3** **4** . Write the chemical formula of the conjugate base of **3** **4** .
4. **1** **2** is the K_b for a weak base. Write the chemical formula of **3** **4** . Write the chemical formula of the conjugate acid of **3** **4** .

5. **1** **2** **3** **4** **5** **6** **7** **8** **9** **0**
6. **1** **2** **3** **4** **5** **6** **7** **8** **9** **0**
7. **1** **2** **3** **4** **5** **6** **7** **8** **9** **0**
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Padhraic Smyth



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