

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) volume, molarity, pH, concentration of ions, amount of substance

(A) $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$

(B) $p = 10^{-\text{pH}}$, $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$, $n = \text{mol}$
 (C) $V = \text{dm}^3$, $p = 10^{-\text{pH}}$, $M = \text{mol dm}^{-3}$, $n = \text{mol}$ (not all the same)

2. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, 10% of the indicator is in its protonated form. What is the pH of the solution? (correct)

(A) $\text{pH} = 4.5 + \log 10$ (B) $\text{pH} = 4.5 - \log 10$

(C) $\text{pH} = 4.5$ (D) $\text{pH} = 4.5 + \log 0.1$ (E) $\text{pH} = 4.5 - \log 0.1$

3. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, 10% of the indicator is in its protonated form. What is the concentration of the indicator? (correct)

(A) 0.1 (B) 0.01 (C) 0.001 (D) 0.0001

4. Which indicator is most appropriate for the following reaction? (correct)



(A) 1 (B) 2 (C) 3 (D) 4

5. A solution of acetic acid and its conjugate base is prepared with a molar ratio of 1:1. The pK_a of acetic acid is 4.75. What is the pH of the solution? (correct)



(A) 4.75 (B) 4.75 (C) 5.25 (D) 5.25 (E) 5.5

6. An acid-base equilibrium is shown. What is the pK_a of the acid? (correct)



Multiple Choice Stoichiometry Test

VM Jensen

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Table of Contents Multiple Choice Stoichiometry Test

1. Understanding the eBook Multiple Choice Stoichiometry Test
 - The Rise of Digital Reading Multiple Choice Stoichiometry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Stoichiometry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test

- Personalized Recommendations
 - Multiple Choice Stoichiometry Test User Reviews and Ratings
 - Multiple Choice Stoichiometry Test and Bestseller Lists
5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks
 - Multiple Choice Stoichiometry Test Public Domain eBooks
 - Multiple Choice Stoichiometry Test eBook Subscription Services
 - Multiple Choice Stoichiometry Test Budget-Friendly Options
 6. Navigating Multiple Choice Stoichiometry Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Stoichiometry Test Compatibility with Devices
 - Multiple Choice Stoichiometry Test Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
 - Highlighting and Note-Taking Multiple Choice Stoichiometry Test
 - Interactive Elements Multiple Choice Stoichiometry Test
 8. Staying Engaged with Multiple Choice Stoichiometry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Stoichiometry Test
 9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
 - Setting Reading Goals Multiple Choice Stoichiometry Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test

- Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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